

Last Name: _____ First Name _____ NetID _____
Discussion Section: _____ Discussion TA Name: _____

Instructions—

This is a closed book exam. You have ninety (90) minutes to complete it.

1. Use a #2 pencil. Do not use a mechanical pencil or pen. Darken each circle completely, but stay within the boundary. If you decide to change an answer, erase vigorously; the scanner sometimes registers incompletely erased marks as intended answers; this can adversely affect your grade. Light marks or marks extending outside the circle may be read improperly by the scanner. Be especially careful that your mark covers the **center** of its circle.
2. Print your last name in the **YOUR LAST NAME** boxes on your answer sheet and print the first letter of your first name in the **FIRST NAME INI** box. Mark (as described above) the corresponding circle below each of these letters.
3. Print your NetID in the **NETWORK ID** boxes, and then mark the corresponding circle below each of the letters or numerals. Note that there are different circles for the letter “I” and the numeral “1” and for the letter “O” and the numeral “0”. **Do not** mark the hyphen circle at the bottom of any of these columns.
4. You may find the version of this Exam Booklet at the top of page 2. Mark the version circle in the TEST FORM box near the middle of your answer sheet. **DO THIS NOW!**
5. Stop **now** and double-check that you have bubbled-in all the information requested in 2 through 4 above and that your marks meet the criteria in 1 above. Check that you do not have more than one circle marked in any of the columns.
6. Print your UIN# in the STUDENT NUMBER designated spaces and mark the corresponding circles. You need not write in or mark the circles in the SECTION block.
7. On the **SECTION line**, print your **DISCUSSION SECTION**. (You need not fill in the COURSE or INSTRUCTOR lines.)
8. Sign (**DO NOT PRINT**) your name on the **STUDENT SIGNATURE line**.
Before starting work, check to make sure that your test booklet is complete. You should have 10 numbered pages plus a Formula Sheet at the end.

*Academic Integrity—***Giving assistance to or receiving assistance from another student or using unauthorized materials during a University Examination can be grounds for disciplinary action, up to and including expulsion.**

This Exam Booklet is Version A. Mark the **A** circle in the TEST FORM box near the middle of your answer sheet. **DO THIS NOW!**

Exam Grading Policy—

The exam is worth a total of 99 points, composed of two types of questions.

MC5: *multiple-choice-five-answer questions, each worth 6 points.*

Partial credit will be granted as follows.

- (a) If you mark only one answer and it is the correct answer, you earn **6** points.
- (b) If you mark *two* answers, one of which is the correct answer, you earn **3** points.
- (c) If you mark *three* answers, one of which is the correct answer, you earn **2** points.
- (d) If you mark no answers, or more than *three*, you earn **0** points.

MC3: *multiple-choice-three-answer questions, each worth 3 points.*

No partial credit.

- (a) If you mark only one answer and it is the correct answer, you earn **3** points.
- (b) If you mark a wrong answer or no answers, you earn **0** points.

**Did you bubble in your name, exam version and network-ID?
Check to make sure you have bubbled in all your answers.**

The next two problems are related.

Radon is a radioactive noble gas (atomic weight 0.222 kg/mole) that leaks from the ground and often gets trapped in homes.

1. What is the average translational kinetic energy of a single radon molecule at room temperature (20° C)?

- a. 1.01×10^{-20} J
- b. 6.07×10^{-21} J
- c. 2.76×10^{-22} J
- d. 4.14×10^{-22} J
- e. Cannot be determined with the information given.

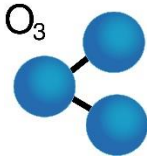
2. Even though it becomes trapped, there is only 10^{-18} kg/m³ of Radon gas at any one time due to its short half-life. What partial pressure does it exert in the home at room temperature?

- a. 1.1×10^{-14} Pa
- b. 1.1×10^{-16} Pa
- c. 2.4×10^{-14} Pa
- d. 2.4×10^{-16} Pa
- e. Cannot be determined with the information given.

3. The noble gases He and Ar can be used to raise and lower the pitch of your voice respectively. What is the approximate ratio of the speed of sound of He to that of Ar, $v_{\text{He}}/v_{\text{Ar}}$?

- a. $\sqrt{1/10}$
- b. $\sqrt{1/2}$
- c. 1
- d. $\sqrt{2}$
- e. $\sqrt{10}$

The next two problems refer to the following situation:



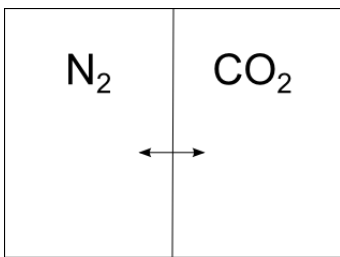
Ozone (O_3) is a non-linear molecule (as shown in figure). Assume the O_3 and O_2 gasses below are at the same temperature and neglect vibrational energy.

4. Which has more total internal energy, 2 moles of O_3 or 3 moles of O_2 ?

- a. 2 moles of O_3
- b. 3 moles of O_2
- c. They contain the same internal energy.

5. Which has a higher translational kinetic energy?

- a. O_2
- b. O_3
- c. They are the same.



6. Two 30-m^3 volumes are separated by a moveable barrier. 1000 kg of N_2 is in the left volume and 1000 kg of CO_2 is in the right volume. Assuming the system is kept at a constant temperature, which way does the barrier start to move when released?

- a. left
- b. right
- c. It doesn't move.

7. What temperature would an ideal gas of hydrogen molecules (H_2) need to be in order for the average speed of each molecule to slow down to 100 m/s?

- a. 0.4 K
- b. 0.8 K
- c. 4 K
- d. 8 K
- e. 40 K

The next two problems refer to the following situation:

Suppose that we are given one mole of a complicated solid with an internal energy per particle of $U=3kT+\eta(kT)^2$ where $\eta=7.4\times 10^{20}$ Joules⁻¹.

8. What is the heat capacity (at constant volume) when the temperature $T=300K$?

- a. 24.9 J/K
- b. 50.4 J/K
- c. 75.9 J/K

9. Suppose that we heat the solid; what can we say about the *change* in internal energy in the following two scenarios: (A) starting from $T=300K$, the temperature of the solid increases by 2K and the volume stays constant, or (B) starting from $T=320K$, the temperature of the solid increases by 2K and the volume stays constant?

- a. The internal energy increases more in scenario A.
- b. The internal energy increases more in scenario B.
- c. The change in internal energy is non-zero and the same in both cases.
- d. The internal energy does not change in either case.
- e. Nothing can be determined from the given information.

The next three problems refer to the following situation:

Suppose that a thin spherical shell of Aluminum (density $2.7 \times 10^3 \text{ kg/m}^3$) with inner radius $R_i=4\text{m}$ and outer radius $R_o=4.01\text{m}$ is immersed in an extremely cold liquid. The shell is filled with a monatomic gas, and the initial temperature of the Aluminum and the contained gas is 293 K. Assume that we can treat the liquid bath as a heat reservoir at temperature 3K.

10. If the thermal conductivity of Aluminum is 200 W/(m-K) , calculate the initial heat-flow, right after the sphere is immersed in the bath.

- a. $1.16 \times 10^9 \text{ W}$
- b. $1.57 \times 10^9 \text{ W}$
- c. 2.90 W

11. Given that the specific heat of Aluminum is 900 J/(kg-K) calculate the heat capacity of the spherical shell.

- a. $2.45 \times 10^6 \text{ J/K}$
- b. $4.90 \times 10^6 \text{ J/K}$
- c. $2.43 \times 10^7 \text{ J/K}$
- d. $4.91 \times 10^8 \text{ J/K}$
- e. $6.56 \times 10^8 \text{ J/K}$

12. Consider the heat transfer process that occurs through the Aluminum between the gas inside the spherical shell and the 3K cold liquid bath outside. Which of the following statements is true?

- (i) The gas inside the spherical shell will continue transferring heat to the liquid and its temperature will asymptote to absolute zero (0 K) after a long time.
- (ii) The rate at which the temperature of the inside gas decreases becomes faster if a denser monatomic gas at the same temperature replaces the initial gas inside the shell.
- (iii) The rate at which the inside gas changes temperature immediately after the shell is immersed doubles if the thickness of the Aluminum shell is cut in half.

- a. i only
- b. ii only
- c. iii only
- d. i and ii
- e. ii and iii

13. Suppose that we have a particle executing random Brownian motion inside the gaseous planet Jupiter. If the particle has an average velocity of 10^6 m/s and an average time between collisions of 10 ms, how long, on average, will it take it to diffuse from the center of the planet to outer space (assume that the distance from the center to outer space is 69,000 km)? Calculate your answer in earth days (1 earth day = 86,400 sec).

- a. 0.53 days
- b. 2.8 days
- c. 8.3 days

The next two problems refer to the following situation:

Consider a spin system of 1,000,000 spins where each spin is equally likely to be up (u) or down (d).

14. The probability that exactly 500,000 spins are up is

- a. greater than 50%
- b. exactly equal to 50%
- c. less than 50%

15. Consider the following two possibilities:

- (1) All the spins are up.
- (2) All odd spins are up and all even spins are down (i.e. ududududu...)

Which of the following statements are true:

- a. (1) and (2) are equally likely
- b. (1) is more likely than (2)
- c. (2) is more likely than (1)

The next three problems are related.

16. Consider two boxes that are separated. Both boxes are 1-m^3 . The first box is filled with 1 mole of N_2 at 100K and the second box is filled with 1 mole of Ar at 200K. The two boxes are placed into contact so heat can transfer between them. What is the final temperature?

- a. 137.5 K
- b. 150.0 K
- c. 162.5 K

17. Now suppose the wall between the two boxes is removed after they have equilibrated. What is the change in the total entropy before and after removing the partition?

- a. 11.5 J/K
- b. 23.1 J/K
- c. 5.76 J/K

18. After the system has equilibrated from the removal of the partition, what is the probability that all the gas is on the left half of the combined box?

- a. $\left(\frac{1}{2}\right)^{1.2 \times 10^{24}}$
- b. $\left(\frac{1}{4}\right)^{3.41 \times 10^{20}}$
- c. $\left(\frac{1}{300}\right)^{2.41 \times 10^{20}}$

The next two questions are related.

On earth at sea level the partial pressure of N_2 is 0.78 atm, at 270 K, but it falls off at higher altitudes. Assume that we can model this using the Boltzmann distribution for a gas in thermal equilibrium.

19. How will the ratio of the partial pressure of N_2 at 1.6 km to the partial pressure of N_2 at sea level change, as the temperature drops to 250 K?

- a. $\frac{p(1.6 \text{ km}, 250 \text{ K})/p(0 \text{ km}, 250 \text{ K})}{p(1.6 \text{ km}, 270 \text{ K})/p(0 \text{ km}, 270 \text{ K})} = 0.98$
- b. $\frac{p(1.6 \text{ km}, 250 \text{ K})/p(0 \text{ km}, 250 \text{ K})}{p(1.6 \text{ km}, 270 \text{ K})/p(0 \text{ km}, 270 \text{ K})} = 1$
- c. $\frac{p(1.6 \text{ km}, 250 \text{ K})/p(0 \text{ km}, 250 \text{ K})}{p(1.6 \text{ km}, 270 \text{ K})/p(0 \text{ km}, 270 \text{ K})} = 1.01$
- d. $\frac{p(1.6 \text{ km}, 250 \text{ K})/p(0 \text{ km}, 250 \text{ K})}{p(1.6 \text{ km}, 270 \text{ K})/p(0 \text{ km}, 270 \text{ K})} = 1.04$
- e. $\frac{p(1.6 \text{ km}, 250 \text{ K})/p(0 \text{ km}, 250 \text{ K})}{p(1.6 \text{ km}, 270 \text{ K})/p(0 \text{ km}, 270 \text{ K})} = 1.06$

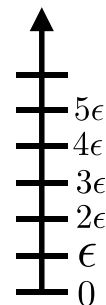
20. At sea level ($h = 0$) the ratio of partial pressures of N_2 and O_2 is $p(N_2, h=0)/p(O_2, h=0) = 0.78/0.21 = 3.7$. At what altitude above sea level will this ratio equal 1? (Hint: sketch the pressure curves as a function of altitude.)

- a. 9.6 km
- b. 75 km
- c. There is no altitude where this occurs.

The next four problems are related.

A vibrational mode of a CO₂ molecule can be modeled as a harmonic oscillator.

21. Assuming the frequency of the oscillation is 7×10^{13} Hz (corresponding to energy levels separated by $\epsilon = 4.6 \times 10^{-20}$ Joules, as illustrated here), what is the highest the temperature can be and still have the probability to be in the vibrational ground state ($E = 0$) greater than 99%?



- 68 K
- 245 K
- 560 K
- 720 K
- Cannot be determined from the information given.

22. Compare the ratio of the probability of having exactly one quantum of energy to zero quantum of energy (i.e., $P(E = \epsilon)/P(E=0)$) vs. the ratio of the probability of having exactly two quantum of energy to one quantum of energy (i.e., $P(E = 2\epsilon)/P(E=\epsilon)$).

- $P(E=\epsilon)/P(E=0) = P(E=2\epsilon)/P(E=\epsilon)$
- $P(E=\epsilon)/P(E=0) > P(E=2\epsilon)/P(E=\epsilon)$
- Whether $P(E=\epsilon)/P(E=0)$ is greater or less than $P(E=2\epsilon)/P(E=\epsilon)$ depends on the temperature.

23. What is the ratio of the heat capacity (from the vibration) of one of these molecules at low temperature ($T \rightarrow 0$) to that at high temperature ($T \rightarrow \infty$)?

- $C(T \rightarrow 0)/C(T \rightarrow \infty) = 0$
- $C(T \rightarrow 0)/C(T \rightarrow \infty) = 1$
- $C(T \rightarrow 0)/C(T \rightarrow \infty) = \infty$

24. Suppose that there were *two* vibrational states with $E = \epsilon$ and one with $E = 0$ (i.e., the first excited level is degenerate). How would the probability $P(E=0)$ to be in the ground state change (compared to the case of a non-degenerate first-excited level, and assuming $T > 0$)?

- $P(E=0)$ would be smaller than if the first excited level had only one state.
- $P(E=0)$ would be the same as if the first excited level had only one state.
- $P(E=0)$ would be greater than if the first excited level had only one state.

Physics 213 Formula Sheet

Spring 2014

Constants, Data, Definitions

$0 \text{ K} = -273.15 \text{ }^\circ\text{C} = -459.67 \text{ }^\circ\text{F}$
 $N_A = 6.022 \times 10^{23} / \text{mole}$
 $k = 1.381 \times 10^{-23} \text{ J/K} = 8.617 \times 10^{-5} \text{ eV/K}$
 $R = kN_A = 8.314 \text{ J/mol}\cdot\text{K} = 8.206 \times 10^{-2} \text{ atm}\cdot\text{mol}\cdot\text{K}$
 $1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$ $1 \text{ liter} = 10^{-3} \text{ m}^3$
 $\text{STP} \rightarrow T = 0^\circ\text{C}; p = 1 \text{ atm}$
 $h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s} = 4.136 \times 10^{-15} \text{ eV}\cdot\text{s}$
 $\hbar = h/2\pi = 1.055 \times 10^{-34} \text{ J}\cdot\text{s}$
 $1 \text{ eV} = 1.602 \times 10^{-19} \text{ J}$
 $c = 2.998 \times 10^8 \text{ m/s}$
 $\mu_e = 9.2848 \times 10^{-24} \text{ J/T}$ $\mu_p = 1.4106 \times 10^{-26} \text{ J/T}$
 $m_e = 9.109 \times 10^{-31} \text{ kg}$ $m_p = 1836 m_e$
 $g = 9.8 \text{ m/s}^2$ $= 1.673 \times 10^{-27} \text{ kg}$

Fundamental Laws/Principles:

First law: $dU = dQ + dW$ Second Law: $d\sigma/dt \geq 0$

$P_i \propto \Omega_i \equiv e^{\sigma_i}$

Classical equipartition $\langle \text{energy} \rangle = \frac{1}{2} kT$ per quadratic term

Entropy & Temperature: $S = k\sigma = k \ln \Omega$; $\frac{1}{T} \equiv \left(\frac{\partial S}{\partial U} \right)_{V,N}$

Heat Capacities: $C_V \equiv (\partial U / \partial T)_V$; $C_p \equiv (\partial(U + pV) / \partial T)_p$

Special properties of α -ideal gases

$U = \alpha NkT = \alpha nRT$ $pV = NkT$ $p_{\text{tot}} = p_1 + p_2 + \dots$

$C_V = \alpha Nk = \alpha nR$ $C_p = C_V + Nk$ $n = \# \text{ moles} = N/N_A$

$c_p/c_v = (\alpha + 1)/\alpha = \gamma$ $W_{\text{by}} = NkT \ln(V_f/V_i)$

$VT^\alpha = \text{const.}$, or $pV^\gamma = \text{const.}$, $\gamma = (\alpha + 1)/\alpha$

$W_{\text{by}} = \alpha Nk (T_1 - T_2) = \alpha (p_1 V_1 - p_2 V_2)$

$\Delta S = C_V \ln(T_f/T_i) + Nk \ln(V_f/V_i)$

α	γ
● 3/2	5/3
●● 5/2	7/5

Processes, Heat Engines, etc

$\Delta U = Q - W_{\text{by}}$ $W_{\text{by}} = \int p dV$

Quasistatic: $dS = dQ/T$ so $\Delta S = \int (C/T) dT$

$dQ = dU + p dV$

$\epsilon_{\text{Carnot}} = 1 - T_C/T_H$

Diffusion and Heat Conduction

$D = (\ell^2/3\tau) = v\ell/3$ $\tau = \ell/v$

$\langle x^2 \rangle = 2Dt$ $\langle r^2 \rangle = 6Dt$

$J_x = \kappa \Delta T / \Delta x$, $\kappa = D_{\text{HC}}$ where $c = C_V/V$

$H_x = J A = \Delta T / R_{\text{th}}$ $R_{\text{th}} = d/\kappa A$ $\Delta L/L = \alpha \Delta T$

$T_A(t) = T_f + (T_{A0} - T_f) e^{-t/\tau}$, $\tau = R_{\text{th}} C_A$

Spins

$\Omega(N, N_{\text{up}}) = \frac{N!}{N_{\text{up}}! N_{\text{down}}!} = \frac{N!}{N_{\text{up}}! (N - N_{\text{up}})!}$; $\Omega(m) = 2^N \sqrt{\frac{2}{\pi N}} e^{-m^2/2N}$; $P(m) = \Omega(m)/2^N$

$M = (N_{\text{up}} - N_{\text{down}}) \mu \equiv m\mu$, $M = N\mu \tanh(\mu B/kT)$

SHO

$P_n = (1 - e^{-\epsilon/kT}) e^{-n\epsilon/kT}$; $\langle E \rangle = \epsilon / (e^{\epsilon/kT} - 1)$ $\epsilon = hf$; $\Omega = \frac{(q + N - 1)!}{q!(N - 1)!}$

Counting, Bin Statistics, Entropy

Ω	Occupancy ($N \ll M$)		
	Unlimited	Single	Dilute
Distinct	M^N	$\frac{M!}{(M-N)!}$	M^N
Identical	$\frac{(N+M-1)!}{N!(M-1)!}$	$\frac{M!}{(M-N)!N!}$	$\frac{M^N}{N!}$

$\ln N! \approx N \ln N - N$

Equilibrium

Boltzmann: $P_n = \frac{d_n e^{-E_n/kT}}{Z}$; $Z \equiv \sum_i d_i e^{-E_i/kT}$

Free energies: $F \equiv U - TS$ $G \equiv U - TS + pV$

Chemical potential:

$\mu = \left(\frac{\partial F}{\partial N} \right)_{V,T} = \left(\frac{\partial G}{\partial N} \right)_{p,T}$; equilibrium $\sum_i (\Delta N_i) \mu_i = 0$

$\mu_i = kT \ln(n_i/n_{Ti}) - \Delta_i$ (ideal gas)

$n_Q = (2\pi mkT/h^2)^{3/2} = (10^{30} \text{ m}^{-3}) (m/m_p)^{3/2} (T/300\text{K})^{3/2}$

Semiconductors $n_e n_h = n_i^2$; $n_i = n_Q e^{-\Delta/2kT}$

Thermal Radiation

$J = \sigma_B T^4$, $\sigma_B = 5.670 \times 10^{-8} \text{ W/m}^2 \text{ K}^4$ $\lambda_{\text{max}} T = 0.0029 \text{ m}\cdot\text{K}$

Particle	mass/mol
N ₂	28g
O ₂	32g
He	4g
Ar	40g
CO ₂	44g
H ₂	2g
Si	28g
Ge	73g
Cu	64g
Al	28g
1g	10^{-3} kg

